Quiz 3A

1. Give the symbol that has the (3 points):
	1. largest atomic radius in period 5 \_\_\_\_\_Rb\_\_\_\_
	2. noble gas with the highest ionization energy \_\_\_\_\_He\_\_\_\_
	3. lowest metallic character energy in period 2 \_\_\_\_\_Ne\_\_\_\_
2. Complete the following sentences with the terms s orbital, s sublevel, p sublevel, p orbital, d sublevel, or f sublevel (4 points). Not all terms are used.
	1. P sublevel is an electron sublevel with 3 orbitals.
	2. D sublevel is an electron sublevel with 5 orbitals.
	3. S orbital is a spherically shaped orbital.
	4. S sublevel is an electronic sublevel that holds up to 2 electrons.
3. How many moles of H are in 6.920 mole of ethanol, C2H5OH (5 points)?

$$6.920 mol C\_{2}H\_{5}OH×\frac{6 mol H}{1 mol C\_{2}H\_{5}OH }=41.52 mol H$$

1. Answer the following questions about the electron configuration (5 points):

1s22s22p63s23p5

* 1. Write the abbreviated electron configuration. [Ar] 3s23p5
	2. Write the electron dot symbol. 
	3. How many core electrons are there? 10
	4. How many valence electrons? 7
1. How long should you rinse your hand if you spill a small amount of chemicals on your hand (2 points)?

You should rinse for at least 15 minutes.

Quiz 3B

1. How many moles of H are in 5.430 moles of acetic acid, CH3COOHHHH (5 points)?

$$5.430 mol CH\_{3}COOH×\frac{4 mol H}{1 mol CH\_{3}COOH }=21.72 mol H$$

1. Give the symbol that has the (3 points):
	1. halogen with the highest ionization energy \_\_\_\_\_F\_\_\_\_\_
	2. period 4 element with the smallest atomic radius \_\_\_\_\_Kr\_\_\_\_
	3. lowest metallic character in alkaline earth metals \_\_\_\_Be\_\_\_\_\_
2. How long should you rinse your hand if you spill a small amount of chemicals on your hand (2 points)?

You should rinse for at least 15 minutes.

1. Complete the following sentences with the terms s orbital, s sublevel, p sublevel, p orbital, d sublevel, or f sublevel (4 points). Note not all terms are used.
	1. F sublevel is an electron sublevel with 7 orbitals.
	2. D sublevel is an electronic sublevel that holds up to 10 electrons.
	3. S orbital is a spherically shaped orbital.
	4. P orbital has two lobes.
2. Answer the following questions about the electron configuration (5 points):

[Ne] 3s23p5

* 1. Write the complete electron configuration. 1s22s22p63s23p5
	2. How many core electrons are there? 10
	3. How many valence electrons? 7
	4. Write the electron dot symbol. 